

18(7), 25(1)

AUTHORS: Krzanowski, Andrzej, and Stegenza, Boleslaw, Engineers

† POL/43-59-4-3/23

TITLE: Carbon and Alloy Tool-Steels

PERIODICAL: Wiadomosci Hutnicze, 1959, Nr 4, pp 114-117 (Poland)

ABSTRACT: The article describes the characteristics of carbonated and casting tool-steels for cold and hot working. The carbonated tool-steels have most simple chemical structures. For high-quality tools, steel containing an appropriate amount of carbonate, sulphur and manganese is used. For tools that have to be ductile to a certain degree, steel Standards N6E and N86 are taken; for wear-resisting tools the carbonated, plain tempered steel of the Standards N9E, N11E, N12E and N13E are used. The tempering capacity of the tool-steel is tested by samples in the size of 25 x 25 mm by heat treatment at temperatures of 760°, 800°, 840° and 880°C. (Table 1). Tool-steels for casting are ✓


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Carbon and Alloy Tool-Steels

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classified into (1) steels for cold working (NC10, NC11, NC6, NWC, NCWV) and (2) steels for warm working. The main components of tool-steels for casting are chromium, tungsten, molybdenum and vanadium. Table 2 shows the most frequent defects of tools occurring during heat treatment. There are 2 tables.

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POL/43-59-5-5/28

18(5.7)

AUTHOR:

Krzanowski, A., and Stegenda, B., Engineers

TITLE:

High Speed Steels

PERIODICAL:

Wiadomosci Hutnicze, 1959, Nr 5, pp 151-154 (Poland)

ABSTRACT:


This article describes in detail the working of high speed steels, their plastic hot and cold transformation, their forging, annealing, hardening and tempering. The main quality of high speed steels is that the cutting tools, manufactured out of them, keep their hardness to a temperature of 600°C, maintaining an edge, even though red hot. Hence, they are called "high speed" or "quick cutting" steels. The elements employed to develop the properties of high speed steels are: tungsten, chromium, molybdenum vanadium and cobalt. To develop the most desirable properties in the steel, the segregates must be uniformly distributed and well scattered and careful control must be kept of melting and casting practice and subsequent forging and rolling operations. Heating for forging

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### High Speed Steels

should be done slowly and gradually, for if a cold bar is placed in a hot furnace, it is liable to be overstrained by unequalized expansion. After forging and machining, high speed steel should be annealed before it is hardened. The temperature of high speed steels must neutralize the inner tension of the steel structure. It is based on the heating of the hardened steel to a temperature of 560-600°C (the temperature itself depends on the kind of steel and the form of the tool) and cooled very slowly in calm air. This is to be repeated several times. The hardness, developed by this operation, is known as secondary hardness. Good results have been obtained by tempering in steam, at a temperature of 300-500°C, in 20 minutes. The tools, thus tempered, are twice as sharp and have a better look. On the surface, there is a thin layer of  $Fe_3O_4$ , which gives the tools an esthetic looking blue hue. Due to the bad thermal conductivity of the high speed steels, the cutting of the tools has to be carried out very cautiously, and absolute cleanness of the wheels has to be maintained.



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### High Speed Steels

In order to improve the cutting efficiency of the tools, they have to be treated 20-30 minutes in a sodium-cyanide bath, the temperature of which is lower than the tempering temperature, by 10-20°C. There are 2 graphs and 1 table.

✓

Card 3/3

STEGENTA, Boleslaw, mgr inz.

The Leipzig Spring Fair in 1959. Wiad hut 15 no.5:158-161  
My '59.

KRZANOWSKI, Andrzej, mgr inz.; STEGENTA, Boleslaw, mgr inz.

Production materials for durable magnets. Wiad hut 15  
no.7/8:218-220 J1-Ag '59.

STEGENTA, Boleslaw, mgr inz.

Conservation, packing, and delivery of drawn materials,  
the priority problem for producer and receiver. Wiad hut  
15 no.10:318 0 '59.



STEGENTA, Boleslaw, mgr., inz.; ZAJC SZ, Hugon, inz.

Spring steel. Wlad hut 18 no.1:12-16 '62.

STEGER, Ferenc, okleveles gepeszmernok, vezeto-tervezo

Endurance tests for railroad diesel engines. Jarmu mezo gep 10 no.  
10:361-367 0 '63.

1. Ganz-Mavag Motorgyartas.

STEGER, Tibor

An illustrated account of the Telephone Factory exhibition.  
Radiotechnika 15 no.6:3 of cover Je '65.

YUGOSLAVIA/Organic Chemistry. Synthetic Organic Chemistry.

G-2

Abs Jour: Ref Zhur-Khim., No 24, 1958, 81714.

Author : Verkade P., Stegerhoek L., Mostert-Pzn S.

Inst :

Title : The Utilization of Silver Salts of Phenylbenzyl  
Phosphoric Acid for the Synthesis of the Monophenyl  
Ester of Phosphatides. (Previous Communication).

Orig Pub: Croat chem acta, 1957, 29, No 3-4, 413-517.

Abstract: The preparation of  $\text{ROP(O)(CH)(OC}_6\text{H}_5\text{)} (I)$  here and  
later, of  $\text{R} = \text{CH}_2\text{CH}_2\text{OCC}_6\text{H}_4\text{H}_3$  is described. From  
 $(\text{C}_6\text{H}_5\text{CH}_2\text{O})_2\text{P(O)}$  and  $\text{SO}_2\text{Cl}_2$  -  $(\text{C}_6\text{H}_5\text{CH}_2\text{O})\text{POCl}$  is  
synthesized from which by the reaction with  $\text{C}_6\text{H}_5\text{ONa}$ ,  
 $(\text{C}_6\text{H}_5\text{CH}_2\text{O})_2\text{P(O)(OC}_6\text{H}_5\text{)}$  was obtained, which by boiling  
with  $\text{NaI}$  in acetone gives the salt  $(\text{C}_6\text{H}_5\text{CH}_2\text{O})(\text{C}_6\text{H}_5\text{O})$

Card : 1/3

STEGINSKY, Bernard

POLAND

STEGINSKY, Bernard

Stello Memorial Institute, Columbus, Ohio

Review, Posteri Fizyki, No 4, 1963, pp 415-25.

"Dynamics of Ionized Gases".

STEGLIK, Josef., master radiolyubitel'skogo sporta.

Czechoslovak radio amateurs. Radio no.10:17 0'55. (MLRA 9:1)  
(Czechoslovakia--Amateur radio stations)

STEGLIK, V.

SILIN, P. professor

"Fundamentals of Sugar Production". M. Drakhovskaya, V. Steglik, and K. Shandera. Reviewed by Silin. Sakh. prom. 30 no 5:78-79 iy '56  
(Sugar industry) (Drakhovskaya, M.) (Steglik, V. Shandrra, K.) (MLRA 9:9)





STEGMAN, KAZIMIERZ.

STEGMAN, KAZIMIERZ. Kultura stawu rybnego. Warszawa, Panstwowe Wydawn.  
Rolnicze i Lesne, 1952. 87 p. (Pondfish culture) DA Not in DLC

AGRICULTURE  
Poland

So: East European Accession, Vol. 6, No. 5, May 1957

STEGMAN, Kazimierz

Selection of carp spawners on Poland's fish farms. Postepy  
nauk roln 10 no.4:109-113 J1-Ag '63.

1. Katedra Rybactwa, Szkoła Główna Gospodarstwa Wiejskiego,  
Warszawa.

STEGMANN, G.

"Principle and Control of Flow Meters Working on Differential Manometers", P. 267. (TECHNICKA PRACA, Vol. 6, No. 5, May 1954, Bratislava, Czechoslovakia)

SO: Monthly List of East European Accessions, (EEAL), LC, Vol. 4, No. 1, Jan. 1955, Uncl.

STEGMANN, G.

Production of half cellulose

p. 63

Vol. 10, no. 4, Apr. 1955

PAPIR A CELULOZA

Praha, Czechoslovakia

SO: Monthly List of East European Accessions, (EEAL), LC, Vol. 5, no.2  
February 1956, Uncl.

STEWART, G.

"Some figures on the Finnish cellulose-paper industry."

PUBLISHED IN: Praha, Czechoslovakia. Vol. 10, no. 8, Aug. 1955.

Monthly list of East European Accessions (EAL), LC. Vol. 8, No. 6, Jun 59, Unclas

STYRMANN, G.

"New automatic mobile irrigation unit, a new technique in plant production."

P. (4) of cover. (Ministerstvo zemědělství --Praha, Czechoslovakia.) Vol. 8, no. 1,  
Jan. 1958

SO: Monthly Index of East European Accession (EEAI) LC, Vol. 7, No. 5, May 1958

L 13312-63

EPF(n)-2/EWT(m)/BDS AFFTC/ASD/AFNL/SSD Pu-4

G/025/63/006/003/006/007

60  
59

AUTHOR: Siewert, G., and Stegmenn, H.

TITLE: Approximating calculations for an irradiation installation with simplified radiation shielding /f

PERIODICAL: Kernenergie, v. 6, no 3, Mar 1962, 124-130

TEXT: The calculations presented demonstrated that it is possible to approximate by relatively simple means the strength of radiation field surrounding an irradiation installation of known dimensions if the composition of the materials used is known. The calculated results correlated satisfactorily with data obtained by actual measurements around an installation housing a 2000-curie cobalt-60 source (which was used as an example to illustrate the calculating procedures involved). The individual steps in the calculations include the determination of the locations for which the prevailing dosage should be calculated; estimation of direct radiation penetrating the roof; estimation of the radiation reflected from the floor and the walls; calculation of total flux, its apparent activity, and scattered radiation at the roof-top; estimation of weakening in radiation through the roof; and calculation of the scattered radiation in the area surrounding the entire installations.

Card 1/2/

Association: People-Owned Enterprise for the Development and Planning of Nuclear-Technological Installations.

STEGMANN, Gustav, doc., inz.

Determination of the characteristic dimension of segmental  
orifice plates. Energetika Cz 13 no.5:240-243 My '63.

1. Slovenska vysoka skola technicka, Bratislava.



STEGMANN, Gustav, doc., inz.

Economic importance of regenerative boilers and summary of the  
experience acquired in Czechoslovakia and abroad. Energetika  
Cz 13 no.6:303-306 Je '63.

1. Slovenska vysoka skola technicka, Bratislava.

STEGMANN, Gustav, doc., inz.

Use of sulfite waste liquor from cellulose production as  
fuel. Tech praca 15 no.11:888-891 N'63.

1. Slovenska vysoka skola technicka, Bratislava.

STEGMANN, Gustav, doc., inz.

Flow measurement of liquids containing sedimentation sludge. Papir  
a celuloza 18 no.4:89-90 Ap '63.

1. Slovenska vysoka skola technicka, Bratislava.

STEGMANN, Gustav, doc. inz.

Methods of calculating universal combustion triangles. Energetika  
Cz 14 no.8:381-385 Ag '64

1. Slovak Higher School of Technology, Bratislava.

L 31752-66 EWP(k)/T-2/EWP(w)/EWP(v) IJP(c) WW/EM

ACC NR: AP6021663

SOURCE CODE: CZ/0032/65/015/008/0575/0581

AUTHOR: Stegmann, G. (Doctor; Engineer)

ORG: Slovak Institute of Technology, Bratislava (Slovenska vysoka skola technicka)

TITLE: Calculation of the optimum arrangement of nozzles for boilers operating on waste sulfate liquor

SOURCE: Strojirenstvi, v. 15, no. 8, 1965, 575-581

TOPIC TAGS: steam boiler, nozzle design, fuel nozzle, combustion

ABSTRACT: The article briefly outlines specific phenomena characterizing the combustion process in steam boilers operating on black sulfate liquor. Then with a simplified physical model a determination is made of the optimum height at which the injecting nozzles should be located. Several simplifying assumptions must be made on account of the complexity of the system, but nevertheless the conclusions indicate ways to improve the efficiency of regenerating boilers. The article was presented by Engineer J. Votoupal. Orig. art. has: 3 figures and 28 formulas. [JPRS]

SUB CODE: 13 / SUBM DATE: none / ORIG REF: 004 / SOV REF: 004

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Card 1/1

UDC: 621.181.6:621.181.67/.68:621.18.001

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VARDAY, Gyorgy, dr.; BICZOK, Imre; OCSVAR, Rezső; LAJTHOS, Zoltan; SZIMELY, Karoly; BERENYI, Akos, dr.; FEHER, Gyula; GALLI, Laszlo; BAKOS, Laszlo; CZIGLINA, Vilmos; GABOS, Gyorgy; SZILAGYI, Gyula; RONAI, Andras; KOVACS, Gyorgy; BACHMANN, Alfred; STEGMULLER, Jozsef; RETHATI, Laszlo; NAGY, Zoltan.

Hydrological questions of the construction industry in Hungary.  
Hidrologiai kozlony 36 no.3:169-170 Je'56.

1. "Hidrologiai Kozlony" szerkeszto bizottsagi tagja (for Galli).
2. "Hidrologiai Kozlony" felelos szerkesztoje (for Kovacs).

STEGNA, G.

"Results of the Mobilization of Plant Power Stations and the Regulation of the Power Supplied from the Network of the Ministry of Fuel and Power Industry to the Lumber, Pulp, and Paper Industries, in Slovakia," p. 9.  
(Energetika, Vol.3, No.1, Jan. 1953, Praha.)

SO: Monthly List of East European Accessions, Vol.2, No.9, Library of Congress, September 1953, Uncl.

STEGNAR, Marjan

~~Prevention of rickets. Zdrav.vest., Ljubljana 24 no.3:91-93 1955.~~  
Prevention of rickets. Zdrav.vest., Ljubljana 24 no.3:91-93 1955.

1. Otroški odderek splosne bolnisnice v Celju - sef Dr. Marjan Stegnar.

(RICKETS, prevention and control)



STEGNER A.

307-RML

✓ 890. AN ANALYSIS OF A  $\Lambda^0$ -PARTICLE (DECAY) IN FLIGHT RECORDED IN NUCLEAR EMULSION. P. Clok and A. Stegner.

539.18

Acta phys. Polon., Vol. 14, No. 1-2, 153-5 (1955).

A  $\Lambda^0$ -particle decay is analysed to obtain an accurate Q-value. The  $\pi^-$ -meson was stopped and its energy therefore known. The momentum of the proton was determined from an analysis of two elastic p-p collisions. The final result is  $36.2 \pm 1.2$  MeV.

D. J. Prowse

①

RML  
for

VAL'TER, A.K.; KOPANETS, Ye.G.; L'VOV, A.N.; STEGNER, A.; TSYTKO, S.P.

Study of the reaction  $Mg^{26}(p, \gamma) Al^{27}$  at proton energies ranging from 1.8 to 2 Mev. Izv. AN SSSR. Ser. fiz. 27 no.11:1419-1426 N '63. (MIRA 16:11)

1. Fiziko-tekhnicheskii institut AN UkrSSR. 2. Institut yadernykh issledovaniy, Varshava, Pol'skaya Narodnaya Respublika (for Stegner).

STECHER, G. Cand Chem Sci -- (diss) "Mechanism and kinetics of <sup>carbon</sup>the formation  
~~of coal~~ <sup>under</sup>during the decomposition of alcohols on catalyzers." Mos, 1959. 10 pp  
(Mos Order of Lenin and Order of Labor Red Banner State Univ im M. V. Lomonosov.  
Chair of Organic Catalysis), 150 copies (KL, 45-59, 144)

5.3000

78059  
SOV/62-60-1-5/37

AUTHORS: Stegner, G., Balandin, A. A., Rudenko, A. P.

TITLE: Influence of Different Stages of Polycondensation  
of the Products of Catalytic Decomposition of Ethyl  
Alcohol on the Rate of Carbonization

PERIODICAL: Izvestiya Akademii nauk SSSR. Otdeleniye khimicheskikh  
nauk, 1960, Nr 1, pp 24-30 (USSR)

ABSTRACT: This is a continuation of the author's previous work  
(Izv. AN SSSR, Chem. Ed., 1959, 1896) on the mechanism  
of carbonization in the decomposition of ethyl alcohol  
over copper-silica. Experimental data presented  
in this paper confirm previous conclusions (see above  
reference) concerning the mechanism of carbonization  
which accompanies catalytic decomposition of ethyl  
alcohol. Carbonization is considered to be a multi-  
stage polycondensation of ethyl alcohol and the pro-  
ducts of its catalytic decomposition. The so-called  
low temperature carbonization (below 600°) proceeds

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Influence of Different Stages of  
Polycondensation of the Products of  
Catalytic Decomposition of Ethyl  
Alcohol on the Rate of Carbonization

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through dehydrogenation of ethyl alcohol. Acceleration or slowing down of ethyl alcohol decomposition (dehydrogenation and dehydration) causes a change in the rate of carbonization. There are 4 figures; and 11 references, 1 German, 10 Soviet.

ASSOCIATION: M. V. Lomonosov Moscow State University (Moskovskiy gosudarstvennyy universitet imeni M. V. Lomonosova)

SUBMITTED: May 4, 1958

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S/062/60/000/011/003/016  
B013/B078

AUTHORS: Stegner, G., Rudenko, A. P., Balandin, A. A.

TITLE: Carbon Formation in the Decomposition of Isopropyl Alcohol, n-Hexyl Alcohol, and Cyclohexanol on the Copper - Silica Gel Catalyst

PERIODICAL: Izvestiya Akademii nauk SSSR. Otdeleniye khimicheskikh nauk, 1960, No. 11, pp. 1930 - 1937

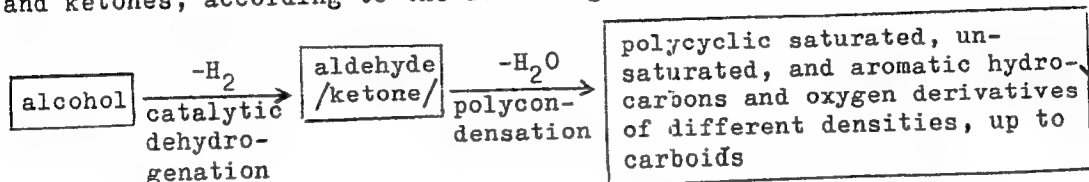
TEXT: A study has been made of the mechanisms of carbon formation in the decomposition of isopropyl and n-hexyl alcohols, as well as of cyclohexanol in the temperature range of 200 - 950°C, proceeding in the same manner as with ethyl alcohol (Refs.1,2). The experiments were conducted in a continuous system for heterogeneous catalytic studies at atmospheric pressure and a volume velocity of 2.25 h<sup>-1</sup>. Fig.1 shows the temperature dependence of the carbon formation rate in the decomposition of the above-mentioned substances. The presence of three mechanisms can be inferred from the course of the curves. Diagrams are suggested for the three mechanisms: the low-temperature mechanism in the temperature

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Carbon Formation in the Decomposition of  
Isopropyl Alcohol, n-Hexyl Alcohol, and  
Cyclohexanol on the Copper - Silica Gel Catalyst

S/062/60/000/011/003/016  
B013/B078

range of 200° - 600°C acts, like a pure polycondensation of aldehydes and ketones, according to the following scheme:

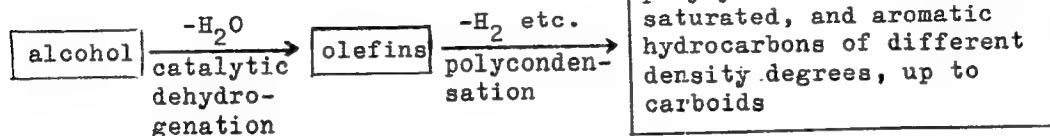


This scheme is applicable to all alcohols concerned, including ethyl alcohol. The intermediate mechanism in the temperature range of 600° - 750°C is a polycondensation of products of the catalytic dehydration of alcohols (propylene, hexylene, cyclohexene). Possibly, a polycondensation of ethylene takes place likewise under these conditions, but since the rate of this process is too low, it occurs only at higher temperatures, in the course of the high-temperature mechanism. The following scheme is offered for the intermediate mechanism:

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Carbon Formation in the Decomposition of  
Isopropyl Alcohol, n-Hexyl Alcohol, and  
Cyclohexanol on the Copper - Silica Gel Catalyst

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B013/B078



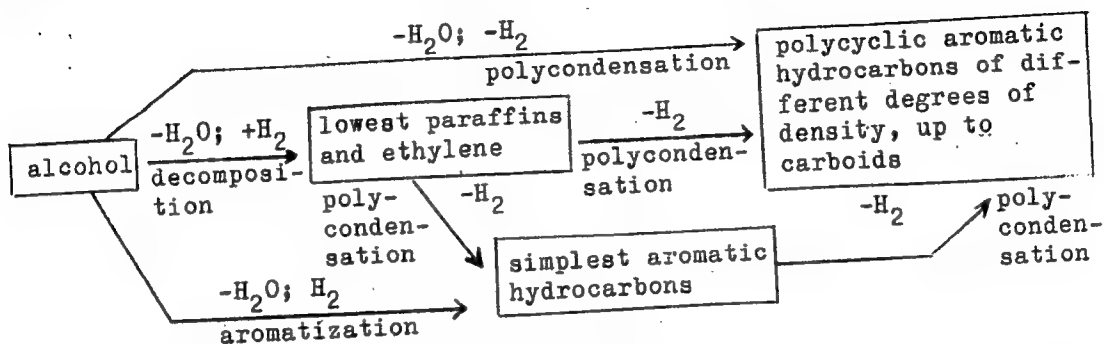
The replacement of the intermediate mechanism by the high-temperature mechanism manifests itself by a marked retardation of the process at temperatures above 750°C. In the range of 750° - 875° the carbon formation is slowed down, after which it is again speeded up vigorously beyond 875°C. This mechanism (750° - 950°C) comprises the following processes: polycondensation of decomposition products of the alcohols used (lowest paraffins and ethylene); polycondensation of aromatic hydrocarbons, resulting from the decomposition of alcohols; polycondensation of the alcohols used. A general scheme applies to them:

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Carbon Formation in the Decomposition of  
Isopropyl Alcohol, n-Hexyl Alcohol, and  
Cyclohexanol on the Copper - Silica Gel Catalyst

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B013/B078



As opposed to the two first-mentioned mechanisms, aromatic hydrocarbons only are given here as the end products. This is explained by the fact that under the conditions of the high-temperature mechanism the formation of saturated, unsaturated, and hydroaromatic polycyclic systems is practically impossible, which is indicated by the composition of

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Carbon Formation in the Decomposition of S/062/60/000/011/003/016  
Isopropyl Alcohol, n-Hexyl Alcohol, and B013/B078  
Cyclohexanol on the Copper - Silica Gel Catalyst

resinous polycondensation products. Their composition and aromatic character is almost the same in all of the alcohols investigated. Table 1 gives the composition of gaseous decomposition products of the alcohols on the copper - silica gel catalyst. The composition of the hydrocarbon part of the gaseous decomposition products of isopropyl and n-hexyl alcohols on the copper - silica gel catalyst is given in Table 2. There are 4 figures, 2 tables, and 4 references: 3 Soviet and 1 German. ✓

ASSOCIATION: Moskovskiy gosudarstvennyy universitet im. M.V. Lomonosova  
(Moscow State University imeni M.V. Lomonosov)

SUBMITTED: June 29, 1959

Card 5/5

BALANDIN, A.A.; RUDENKO, A.P.; STEGNER, G.

Formation of coal dendrites in the course of decomposition of  
alcohols on nickel. Ozv.AN SSSR.Otd.khim.nauk no.5:762-770 My  
'61. (MIRA 14:5)

1. Moskovskiy gosudarstvennyy universitet im. M.V.Lomonosova.  
(Alcohols) (Coal)

STEONER, H.

Acetone from Acetylene and Water, Part II by K. MARCZEWSKA, H. STEONER  
J. MOSCOWICZ and E. TRZESZCZANOWICZ, Page 463, Przemysl Chemiczny, No. 6, 1957.

Stegner, H.

7 7  
Acetone from acetylene and steam. I. K. Marczyńska,  
Musierowicz, H. Stegner, and E. Treszczanowicz [Inst.  
Chem. Ogólnej, Warsaw]. *Przemysł. Chem.* 13, 495-10  
(1957)(English summary).—A pilot-plant installation is  
described for the synthesis of acetone from dil.  $C_2H_2$  based  
on the method used at Oppau, Germany.  $C_2H_2$  is obtained  
from partial burning of methane which is a cheap raw ma-  
terial in Poland. A full-scale plant working on this method  
was recently constructed in Hungary. The empirical equa-  
tion for the synthesis is:  $C_2H_2 + 1.41 H_2O \rightarrow 0.47 CH_3CO-$   
 $CH_3 + 0.06 CH_3CHO + 0.41 CO_2 + 0.07 CO + 0.50 H_2$ .  
The reaction is exothermic and is carried out with 10-fold  
excess of steam at  $450^\circ$ . In order to prevent polymerization  
of  $C_2H_2$  the temp. on the catalyst should not exceed  $500^\circ$ .  
A catalyst was prepd. by mixing 1 part of contact mass used  
for the conversion of water gas with 3 parts of  $ZnO$  (Zn  
white). The compn. of the mixt. was: Zn 80, Fe 14, and  
Cr 1.3% by wt.; it was pelleted to 8 mm. diam.  $\times$  6 mm.,  
bulk d. before activation 1.68, after activation 1.63, and,  
after using them for 50 hrs., 1.99 g./cc. The activation

6  
4E3d

STEGNER, Halina

Obtaining tetrachloroethylene from carbon tetrachloride.  
Przem chem 42 no.6:306-309 Je '63.

1. Zaklad Syntezy Kontaktowej, Instytut Chemii Ogolnej,  
Warszawa.

5. GMR, U.

"Installation of Solar Heating in a Bus", I. BIA, (GOSPODARSTWA ROLSKA,  
T. 1. 4, No. 10, October 1974, Warsaw, Poland)

2. Monthly list of East European Acquisitions (GMR), LG, Vol. 4, No. 3,  
March 1976, incl.

STEGNER, Z., mgr inż.

Current information concerning patents. Techn motor 12  
no. 4/5: 144-146 Ap-May '62.

1. Biuro Instrukcyjne Przemysłu Motoryzacyjnego, Warszawa.



NAVAKATIKYAN, A. O., kand. med. nauk; LYUBOMUDROV, V. Ye., kand. med. nauk; SHCHERBAKOVA, O. I.; PAVLOVA, O. A.; BASAMYGINA, L. Ya.; STEGNIY, A. S. (Donetsk) U.S.S.R.

Evaluation of the arterial pressure in workers of certain professions. Vrach. delo no.7:136 J1 '62. (MIRA 15:7)

1. Laboratoriya klinicheskoy fiziologii (zav. - kand. med. nauk A. O. Navakatikyan) i otdel professional'nykh zabolevaniy (ispolnyayushchiy obyazannosti zaveduyushchego - kand. med. nauk V. O. Lyubomudrov) instituta fiziologii truda i kafedra fakul'tetskoy terapii II (zav. - dotsent N. S. Kamenetskiy) meditsinskogo instituta.

(BLOOD PRESSURE)

STEGNIY, A. Ye.

Stegniy, A. Ye. - "On the problem of physiotherapy in the treatment of ulcers from stumps," Uchen. zapiski (Ukr. nauch.-issled. in-t protezirovaniya), Issue 1, 1948, p. 79-90, - Bibliog: 11 items

SO: U-4355, 14 August 53, (Letopis 'Zhurnal 'nykh Statey, No. 15, 1949)



STEGR, Jar.

STEGR, Jar., Dent. (Praha)

Technic of roentgenography. Prakt. sub. lek., Praha 2 no.4:  
83-94 1954.

(TEETH, radiography,  
technic)

STEGUMIN, S.I.

Courses for public health organizers in Kuybyshev. Zdrav.Ros.  
Fed. 2 no.9:44-45 S'58 (MIRA 11:10)  
(KUYBYSHEV---PUBLIC HEALTH)

STOLIK, S.I., Phil's Sci — (Mss) <sup>7/1</sup> "History of ~~the~~ Kuybyshev Medical  
Institute in connection with the development of higher medical education  
in the USSR." Kuybyshev, 1959. 38 pp (Kuybyshev State Med Inst. Chair  
of Organization of Public Health and History of Medicine). 250 copies  
List of author's works pp 37<sup>38</sup> (10 titles) (p1, 40-59, 106)

YEROSHEVSKIY, T.I., prof.; STEGUNIN, S.I., assistant

Kuybyshev Medical Institute during the years of Soviet power. Trudy  
Kuib.med.inst. 11:3-12 '60. (MIRA 15:8)

1. Kafedra organizatsii zdravookhraneniya i istorii meditsiny  
Kuybyshevskogo meditsinskogo instituta (for Stegunin).  
(KUYBYSHEV--MEDICAL COLLEGES)

STEGUNIN, S.I.

Nikolai Alekseevich Anan'ev; on the anniversary of his death. Sov.  
zdrav. 20 no.5:89 '61. (MIRA 14:5)  
(ANAN'EV, NIKOLAI ALEKSEEVICH, 1901-1960)



ARTEMOV, P.I.; ZINOV'YEV, G.A.; STEGUNIN, S.I. (Kuybyshev)

Diseases of the circulatory organs among the population of  
Kuibyshev (as revealed by extensive study of disease incidence  
in 1958). Sov.zdrav. 21 no.8:47-50 '62. (MIRA 15:11)

1. Iz kafedry organizatsii zdravookhraneniya i istorii meditsiny  
(zav. - prof. I.M.Bulayev) Kuybyshevskogo meditsinskogo instituta.  
(KUYBYSHEV--CARDIOVASCULAR SYSTEM--DISEASES)

ARTEMOV, P.I.; STEGUNIN, S.I.

Basic problems of dispensary follow-up of patients. Sov. med. 28  
no. 4:127-129 Ap '64. (MIRA 17:12)

1. Kafedra organizatsii zdravookhraneniya i istorii meditsiny  
(zav. - dotsent S.I. Stegunin) Kuybyshevskogo meditsinskogo  
instituta.

SEBENDA, J.; STEHLICEK, J.;

Alkaline polymerization of 6-caprolactams. Pt.10. Coll  
Cz Chem 28 no.10:2731-2743 O '63.

STEHLICEK, J.; SEBENDA, J.; WICHMAREK, V.

Alkaline polymerization of  $\epsilon$ -caprolactam. Pt.17. Coll Cz Chem 29  
no.5:1236-1258 Py '64.

1. Institute of Macromolecular Chemistry, Czechoslovak Academy  
of Sciences, Prague.

CZECHOSLOVAKIA

STEHLICEK, J; GEHRLE, K; SEBENDA, J

Institute of Macromolecular Chemistry, Czechoslovak  
Academy of Sciences, Prague - (for all)

Prague, Collection of Czechoslovak Chemical Communi-  
cations, No 1, January 1967, pp 370-381

"Alkaline polymerization of 6-caprolactam. Part 26:  
N-carbamoylcaprolactams as activators of the alkaline  
polymerization of caprolactam."

CZECHOSLOVAKIA

STEHLICEK, J; LABSKY, J; SEBENDA, J

Institute of Macromolecular Chemistry, Czechoslovak  
Academy of Sciences, Prague - (for all)

Prague, Collection of Czechoslovak Chemical Communi-  
cations, No 2, February 1967, pp 545-557

"Alkaline polymerization of 6-caprolactam. Part 25:  
The effect of structure of the acyl on polymerization  
activated by acylcaprolactams or diacylamines."

# CZECH

A method for the determination of organically bound  $\text{SO}_2$  in sulfonated oils. Antonín Stehlík and Boleslav Nováček (Leather & Allied Trades Research Inst., Otrokovice, Czech.). *Českoslov. kovářství* 3; 56-7(1953).—Known methods (Hart, C.A. 28, 4250<sup>4</sup>, Burton and Byrne, C.A. 41, 1117<sup>1</sup>) were tested. For isolation of sulfo esters: shake a weighed 1.5-3.0-g. sample with 20 ml.  $\text{Et}_2\text{O}$  and 20 ml. satd. NaCl soln. contg. 12 g. AcONa and 5 ml. AcOH per 250 ml. satd. NaCl soln. Drain, wash the ether layer with 20 ml. satd. NaCl soln., and evap. the ether. Add 20 ml. 0.1N NaOH and  $\text{H}_2\text{O}$ , evap., ignite for 15 min. with a small flame, wet with  $\text{H}_2\text{O}_2$  (10%), evap., and ash at 500-60°. Titrate the ash against methyl red and methylene blue with 0.05N  $\text{H}_2\text{SO}_4$ . A crucible of stainless steel, but not of Pt, can be used. When only  $\text{CHOSO}_2\text{H}$  groups are present, the results correspond to those of the gravimetric method. If  $\text{CHSO}_2\text{H}$  groups are present, the results by the new method are higher, the increase corresponding to  $\text{CHSO}_2\text{H}$  groups present; these can be calcd. from the difference. L. M.

Stehlík, Antonín

# CZECH

The absorption of water by leather. Antonín Stehlík  
(Leather & Allied Trades Research Inst., ~~Cottbus~~  
Czech.). *Českoslov. kůžedářství* 3, 84-5 (1953).—The Kubelka-  
Němec method (C.A. 32, 61974) of measuring the water  
absorption of leather is exact, but a considerable error may  
occur, if the H<sub>2</sub>O content of the tested leather is not con-  
sidered. The author recommends calcg. the H<sub>2</sub>O absorption  
to 11% initial H<sub>2</sub>O content. L. Masner



STEFELIK, J. P. L. N. V.

# CZECH

✓ Relation between the shake and filter bell methods of tannin analysis. Antonín Stehlík, M. Gošová, and B. Lačňáková (Leather Research Inst., Gottwaldov, Czech.). *Kotafisť* 4, 105-7(1954).—The results can be calcd. according to the equation  $T' = T_1 + N_1[1 - (N/N_1)]$ , in which  $T$  = tannin by the filter bell (I) method of analysis,  $N$  = nontannins by I,  $T_1$  = tannin by the shake method (II), and  $N_1$  = nontannins by II. The value  $N/N_1$ , characteristic for each tanning material, must be known and be const. These values have been detd. by the Czechoslovakian standard hide powder. Values  $K = 1 - N/N_1$  and  $K_1 = N_1/N$  have been detd. For quebracho ordinary, quebracho sulfited, Syntar K.D, Syntan SN<sub>11</sub>, Kortan QU 1, Svitay (m/n ext.), spruce, chestnut, oakwood exts., and val. 12.  $K$  resp.  $K_1$  is 0.41, 0.39, 0.09, 0.30, 0.37, 0.24, 0.7, 0.44, 1.62, 1.43, 0.39, 0.42, 0.59, 0.37, 0.48, 0.17, 0.7, and 0.29 resp. A table of results of calcs. is given.

Stehlík, Antonín

New analytical methods in tanneries. Antonín Stehlík  
(Leather and Allied Trades Research Inst., Gottwaldov, Czech.). *Kobalství* 5, 35-7(1955).—In testing leather for permeability to air, Podor's app. (from U.S.S.R.) gives more comparable results than does Bergmann's or Schopper's app. A hydrothermal test for the detn. of stability of vegetable-tanned leather has been introduced. A new method for the detn. of organically bound SO<sub>2</sub> in sulfated oils (C.A. 49, 8017a) gives better results if excess H<sub>2</sub>SO<sub>4</sub> is added, CO<sub>2</sub> removed by boiling, and the soln. back-titrated. For detn. of sulfonyl chlorides the sample is burned in the Grote-Krecker (C.A. 27, 1841) app., and the combustion products are absorbed in alk. H<sub>2</sub>O<sub>2</sub> soln. The ratio of Cl to SO<sub>2</sub> is always higher than 1; hence Cl is also bound by addn. or substitution. The SO<sub>2</sub>Cl content calcul. from  $-\text{SO}_2\text{Cl} + 2\text{KOH} = -\text{SO}_3\text{K} + \text{KCl} + \text{H}_2\text{O}$  gives the same results as the polarographic detn. of Majranovskij and Neiman (C.A. 46, 28e). For the polarographic detn. of sulfonyl chloride, cyclohexanone is a suitable solvent; the half-wave potential is near zero (to the calomel electrode). There is a linear relation between the concn. of sulfonyl chloride and the height of the wave, unless the proportion of solvents is changed. L. Master

STEHLIK, A.

70th anniversary of reclamation activity.

p. 213  
Vol. 5, no. 6, June 1955  
VODNI HOSPODARSTVI  
Praha

SO: Monthly List of East European Accessions (EEAL), LC, Vol.5, no. 3  
March 1956

CZECHOSLOVAKIA / Chemical Technology. Chemical Prod- H-35  
ucts and Their Application. Leather.  
Fur. Gelatin. Tanning Agents. In-  
dustrial Proteins.

Abs Jour: Ref Zhur-Khimiya, No 1, 1959, 3364.

Author : Muller, K., Stehlik, A.

Inst : Not given.

Title : The Effect of Tanning Agents Blends Which Are  
Used in Czechoslovakia Upon the Tanning of Lower  
Part of Shoe Wear, and the Application of Some  
Less Frequently Used Methods For Their Determin-  
ation.

Orig Pub: Veda a vyzk. v prumyslu kozedeln., 1956, 1, 7-26.

Abstract: The properties of natural tanning agents which  
are used in Czechoslovakian leather factories  
were investigated. The content of salts and  
weak acids, the degree of astringency were de-

Card 1/2

CZECHOSLOVAKIA/Chemical Technology - Chemical Products and  
Their Application. Leather. Mechanical Gelatins.  
Tanning Materials. Technical Albumins.

H-35

Abs Jour : Ref Zhur - Khimiya, No 17, 1958, 59699

Author : Stehlik Antonin, Kokes Drahoslav, Podolska Miroslava,  
Muck Eduard

Inst : -

Title : Determination of the Degree of Sulfitization of  
Vegetable Tanning Materials.

Orig Pub : Veda a vyzk. v prumyslu kozedeln., 1956, 2, 53-58

Abstract : The method is based on the determination:  
1) of the quantity of free  $\text{SO}_2$  (from sulfite or bi-  
sulfite) and  
2) of the quantity of  $\text{SO}_2$  connected with the tanning ma-  
terial. The method provides reproducible results and  
can be applied in any laboratory.

Card 1/1

- 115 -

Stehlík, A.

✓ Direct volumetric determination of soluble sulfides. A.  
Stehlík and M. Ambrož (Výzkumný ústav kožedělný, Olom-  
bovice, Czech.). Chem. Listy 50, 1320-1 (1956). S<sup>2-</sup> is  
detd. besides SO<sub>3</sub><sup>2-</sup> and S<sub>2</sub>O<sub>3</sub><sup>2-</sup> by complexometric titration  
with 0.05N ZnSO<sub>4</sub> and Eriochrome Black T (preferably A  
or B) as indicator. Max. deviations from the mean value  
were +1.6 and -1.9%. Fe<sup>++</sup> and Mg<sup>++</sup> interfere in amts.  
exceeding 3 and 0.2 mg., resp. L. J. Urbánek

Chem 2

600

**"APPROVED FOR RELEASE: 08/25/2000**

**CIA-RDP86-00513R001653110011-2**

**APPROVED FOR RELEASE: 08/25/2000**

**CIA-RDP86-00513R001653110011-2"**

6  
Melting point of pure tellurium. A. SIMKE AND B. STREHLER. *Collection Czechoslov. Chem. Communications* 2, 303-14(1930).—Com. Te was purified until spectroscopically pure. Heating curves showed no crit. point from room temp. to m. p. Te m. 482.0° in vacuo. This point is lowered in H<sub>2</sub> and CO<sub>2</sub> by 0.15-0.2°, resp., because of soln. of the gases.  
ANN NICHOLSON HIND

The melting point of tellurium dioxide. A. SIMKE AND B. STREHLER. *Collection Czechoslov. Chem. Comm.* 2, 447-56(1930); cf. *C. A.* 24, 4228. —TeO<sub>2</sub> m. 732.6° ± 0.1°. This value was detd. by a special extrapolation from heating curves. TeO<sub>2</sub> crystallizes from the melt; it is uniaxial and pos.; d. = 6.02.  
V. F. HARRINGTON

AS 51.4 METALLURGICAL LITERATURE CLASSIFICATION



*BC*

*d-1*

**Volcanic decomposition of malachite by oxidation of quinine to semi-volatile malachite by means of some sulphate or potassium persulfate.** — R. S. GILMAN (Chem. Ind., 1962, 80, 455-457). — Malachite was heated to 300°C by means of FeCl<sub>3</sub>, SnCl<sub>4</sub>, PbO<sub>2</sub>, Bi<sub>2</sub>O<sub>3</sub>, Sb<sub>2</sub>O<sub>3</sub>; H<sub>2</sub>O<sub>2</sub> is added and the hot solution is treated with CuSO<sub>4</sub> or KNO<sub>3</sub> in presence of MnCl<sub>2</sub>, and in an atm. of CO<sub>2</sub>, the end point being determined potentiometrically.

R. T.

ATM-SLA METALLURGICAL LITERATURE CLASSIFICATION

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2

A STUDY OF THE PERMEABILITY OF THE MEMBRANE L. D. STOKES. Collection (Zach-  
 elow. Chem. Comm. 9, 434-56(1937)). Curves obtained by  
 means of a Mohr's counter (C.A. 25, 500) indication vol.  
 increase of soln. acid, from the pure solvent: a good membrane  
 unanalyzed theoretically. The extrapolated initial velocity  
 of ascent of the meniscus in the capillary tube of the counter is  
 proportional to the concn. of the soln. and the initial velocity  
 of descent of the meniscus is proportional to the difference in the v.l. of the  
 soln. and the solvent (being proportional to the difference in the v.l. of the  
 soln. and the solvent for both components of the soln.) After cor-  
 recting for the effect of hydrostatic pressure the curves have an  
 exponential form: the retardation of the displacement velocity of  
 the meniscus is caused by gradual diln. of the soln. and depends on  
 the permeability of the membrane for the solvent. When the initial  
 hydrostatic pressure is zero the concn. of the capil. curves is  
 proportional to the concn. of the soln. whereas the time needed for  
 attaining the max. is independent of concn. C. A. R. 1937

ASAC S.A. METALLURGICAL LITERATURE CLASSIFICATION

62

CA

7

Potentiometric determination of  $\text{MoO}_3(\text{OH})$  in 3.5-6.3 N hydrochloric acid at 80°. Blahoslav Stehlik. *Chem. Listy* 38, 13 (1944). —The bluish compl. produced by reduction of Na molybdate with Zn, SnCl<sub>2</sub>, Pb, or TiCl<sub>3</sub> in strongly acidic soln. can be titrated with  $\text{Ce}(\text{SO}_4)_2$  or  $\text{KMnO}_4$  in the presence of  $\text{MnSO}_4$  if protected against air.  
 $\text{MoO}_3(\text{OH}) + \text{Ce}^{4+} + \text{H}_2\text{O} \rightarrow \text{MoO}_4^{2-} + \text{Ce}^{3+} + 3\text{H}^+$   
Milos Hudlicky

CA

2

Antoni Stank. Blahoslav. Stehlik. Chas. Lisy  
to, 22-00(1948).—Obituary. M. Hudlický

CA		10	
<p><b>PROCESSES AND PROPERTIES INDEX</b></p> <p>Molecular compounds of carboxylic acids with monohydric alcohols, indicated by means of the Ubbelohde osmometer. B. Stehlik. <i>Chem. Zvesti</i> 1, 97-110, 120-34 (1947).--Curves connecting velocity of osmosis with mol. compn. of the binary systems exhibit well-defined min., suggesting formation of the mol. compds.: <math>\text{HCO}_2\text{H} \cdot 3\text{MeOH}</math>; <math>\text{AcOH} \cdot 3\text{MeOH}</math>; <math>\text{HOCH}_2\text{CO}_2\text{H} \cdot 4\text{MeOH}</math>; <math>(\text{CO}_2\text{H})_2 \cdot 8\text{BuOH}</math>; <math>(\text{CH}_2\text{CO}_2\text{H})_2 \cdot 8\text{MeOH}</math>; <math>[\text{CH}(\text{OH})\text{CO}_2\text{H}]_2 \cdot 8\text{BuOH}</math>; and citric acid <math>\cdot 10\text{MeOH}</math>. In general, the no. of mols. of alc. in the compds. is equal to the no. of OH groups in the acids, regarding <math>\text{CO}_2\text{H}</math> as <math>\text{C}(\text{OH})_1</math>. B. A.</p>			
<p><b>ASAC-SLA METALLURGICAL LITERATURE CLASSIFICATION</b></p>			
<p><b>INDEXING</b></p>			



PROCESSES AND PROPERTIES INDEX																									
<p>CA 10</p> <p>Coordination of monohydric alcohols, ethers, or acetone to hydrogen ions. B. Stehlik (Slovak. Inst. Technol., Bratislava). <i>Collection Czechoslov. Chem. Commun.</i> 12, 516-21(1947); cf. C.A. 42, 436f.—By measuring the initial rates of osmosis through a rush membrane of aquinats. of 0.25 M H<sub>2</sub>SO<sub>4</sub> or 0.5 M HCl with the following substances, compd. formation is shown to occur between a H ion and 6 mols. MeOH, 4 EtOH, PrOH, iso-PrOH, or Me<sub>2</sub>CO, 3 BuOH, <i>tert</i>-BuOH, or Et<sub>2</sub>O. The max. mol. wt. of the coordinated mols. produced lies between 240.5 and 256.3. The H ion belongs to the pos. group of substances of Kell5 (cf. preceding abstr.) and reacts only with his neg. substances. David Lewis</p>																									
<p>ASB. S.L.A. METALLURGICAL LITERATURE CLASSIFICATION</p>																									
<p>13001 STEHLIK</p>																									



CA

2

Crystal structure of tellurium dioxide. Blahoslav Stehlik and Ladislav Balak. (Chem. Zvesti 2, 6-12, 33-45, 69-79 (1948)). TeO<sub>2</sub> prepd. by crystn. from H<sub>2</sub>SO<sub>4</sub> soln. or by fusion was examd. by a rotating crystal using Cu K $\alpha$  radiation. The tetragonal cell unit contains 4 mols. of these sites:  $a = 4.706 \text{ \AA}$  and  $c = 7.504 \text{ \AA}$ . The space group is  $D_2^2 = P4_2/m2$  or  $D_2^2 = P4_2/m2$ . The parameters of the Te atom are  $x = y = 0.00$ ,  $z = 0$ , and of the O atom are  $x = 0.177$ ,  $y = 0.227$ ,  $z = 0.217$ . The lattice is formed by a three-dimensional covalent chain of atoms.

Jan Mucka

ASD-3LA METALLURGICAL LITERATURE CLASSIFICATION

SEARCHED	INDEXED	SERIALIZED	FILED
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100	100	100	100

-A

The coordination of monohydric alcohols to the hydrogen atom in the hemiacetal hydroxyl. Micholay, Svitlik. *Chem. Zvesti* 2, 79-81 (1948). The coordination has been proven by the measurement of rates of osmosis through the Olehin rush membrane for glucose, fructose, and maltose. The alkylene oxide form of these saccharides was confirmed. Cf. *Chem.* 42, 1525d. For Mols.



*pol*

Constitution of orthotelluric acid. Blahoslav Strblík.  
Chem. Zvesti 2, 220-31(1948).—Mol. compds. with 12  
mols. of BuOH and 18 mols. of PrOH, EtOH, or MeOH are  
indicated by the osmotic studies. Six OH groups in  
 $H_6TeO_6$  are confirmed.

Jan Micka

1st AND 2nd GROUPS

PROCESSES AND PROPERTIES INDEX

6

The principle of the structure of silicates. Vishovsky  
 Stehlik. Chem. Zvesti 2, 235-41(1948). Jan 1948

COMMON ELEMENTS

PERIODIC TABLE

ASB-SLA METALLURGICAL LITERATURE CLASSIFICATION

REGION 1 (1948-1950)

REGION 2 (1951-1955)

REGION 3 (1956-1960)

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1ST AND 2ND ORDERS

PROCESSES AND PROPERTIES - 1001

100 AND 4TH ORDERS

CA

2

An osmotic study of ammonia. Blahoslav, J. *Chem. Zvesti* 2, 261 (1948). The fresh membrane indicates 2 mol. compds. in dil. mixts. of  $\text{NH}_3$  and  $\text{HCl}$ .  $\text{NH}_3$ ,  $\text{HCl}$  and  $4\text{NH}_3$ ,  $\text{HCl}$ . Mol. compds. of  $\text{NH}_3$  with 2 butanols or 3 lower alcs. indicate that  $\text{NH}_3$  is a 11 ion coordinated by one mol. of  $\text{NH}_3$ . Jan Miska

ASAC, S.A. METALLURGICAL LITERATURE CLASSIFICATION

1 2 3 4 5 6 7 8 9 10 11 12 13 14 15 16 17 18 19 20 21 22 23 24 25 26 27 28 29 30 31 32 33 34 35 36 37 38 39 40 41 42 43 44 45 46 47 48 49 50 51 52 53 54 55 56 57 58 59 60 61 62 63 64 65 66 67 68 69 70 71 72 73 74 75 76 77 78 79 80 81 82 83 84 85 86 87 88 89 90 91 92 93 94 95 96 97 98 99 100

100 AND 4TH ORDERS

Ortho effect of nitro phenols. B. Stchlik. *Chem. Listy*  
42, 171-3 (1948). Expts. on a rush-membrane osmometer  
showed that *o*-nitrophenol does not form a mol. compd. with  
alcs. in contrast to phenol, which binds 1 mol. of an alc.  
The different behavior of these 2 phenols is ascribed to the  
intramol. H bond which prevents the phenolic H from form-  
ing a bridge to an alc. HO group. M. Hudlický

CA

Intramolecular bridge in pyrocatechol and hydroquinone. Blahoslav Stehlik. *Chem. Zvesti* 2, 81-4(1948).  
—Mol. compds. of dihydroxybenzenes with monohydric alcs. have been found by osmosis through a rush membrane, showing that there is an intramol. H bridge in pyrocatechol and hydroquinone. Jan Micks.



I. Osmometric study of aniline. Miroslav Stehlik (Tech.

Univ., Bratislava, Czech.). Chem. Zvesti. 23(1979).

By the rush-membrane method mol. compds. of 2 mols. of aniline (I) with 1 mol. of MeOH, butanol, Et<sub>2</sub>O, or acetone are indicated. The N atom of I is connected with the O atom of the other compd. by the H<sub>2</sub>O mol. In mixts. of I with HCl, the anomalous (C<sub>6</sub>H<sub>5</sub>NH<sub>2</sub>)<sub>2</sub>.HCl was ascertained. Its rush no.  $\alpha = 4$  indicates that by the coordination of I to H ion the polarity of amine hydrogens is increased and that at the same time the H ion loses the ability to coordinate aq.

Jan Miska

C/★

10

Mesohydric tautomerism of urea. Blahoslav Stehlik and Alexander Tkáč. *Chem. Zvesti* 3, 33-8 (1949). - According to Hunter's conception of the mesohydric tautomerism (C.I. 40, 1892), by osmosis through a rush membrane, urea has been found to have the structure:  $\text{H-NH-C(NH}_2\text{)-O}$ . In a mixt. with HCl, where the unchelated N is coordinated to the H ion, the anomalous ammonium salt  $[\text{CO-NH}_2]_2\text{HCl}$  has been found.

Jan Micka

CA

Chelation of allcinamic acid. Blahoslav-Stehlik.  
Chem. Zvesti 3, 71-6(1949).—Osmosis through the rush  
membrane points to a mol. compd. of  $\text{PhCH:CHCO}_2\text{H}$   
with 15 mols. butanol, which indicates that the H atom on  
the benzene ring in the ortho position forms a bridge to  
the carbonyl O. Jan Micka

10

**An osmometric study of chloroform and chloral hydrate.**  
Blahoslav Stehlik and Alexander Tkáč (Slovak Tech  
Univ., Bratislava, Czech.). *Chem. Zvesti* 3, 161-68  
(1959).—It was found by the osmometric method with a  
ruth membrane (1) that the H atom of  $\text{CHCl}_3$  is polar and  
able to form a H bridge with the O atoms of aces. and (2)  
that in chloral hydrate the HO groups are closed into 2  
chelated rings by the bridge  $\text{O}-\text{H} \cdots \text{Cl}$ . J. M.

27

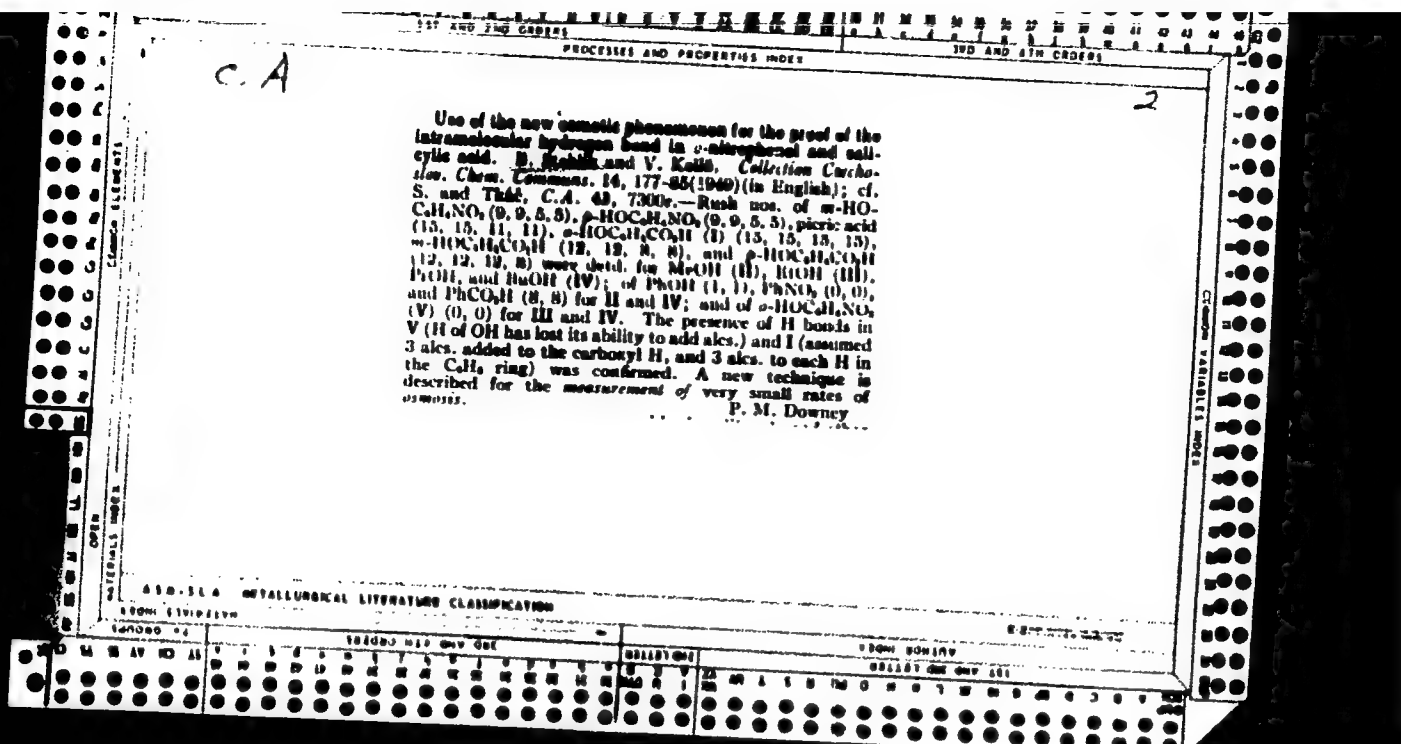
13

An osmometric study of alloxan and dimethylalloxan.  
Machoslav Stehlik (Slovak Tech. Univ., Bratislava, Czech).  
Chem. Zvesti 3, 325-32 (1949). By using a rush membrane  
the mesohydric tautomerism of alloxan has been proven.  
One mol. of H<sub>2</sub>O in alloxan and dimethylalloxan has been proven  
of crystn. but chemically bound. Two mols. dimethyl-  
alloxan in BuOH form with 2 mols. HCl an anomalous salt  
which is decompl. by the excess of BuOH. Jan Miska

CA

New osmotic phenomenon of some acids and sugars.  
H. H. H. and A. H. H. Collection (Carbide). Chem.  
Comm. 14, 10-19(1949)(in English). To explain the  
osmotic phenomena observed previously (C.A. 42,  
430f) and with some acids and sugars, the rank number,  $r$ ,  
is introduced, which represents the no. of mols. of alc.  
combined with one mol. of another compd. through co-  
ordination with the H of the alc., and is represented by  
the equation  $x = \sum h_i (h_i - 1)$ , where  $h_1, h_2, \dots, h_n$  are the  
no. of H in the functional groups 1, 2, ...,  $n$  and  $h$  is the co-  
ordination no. of the H in 1, 2, ...,  $n$ . The Ubbelohde osmome-  
ter was used; a cylindrical rank membrane was filled  
with mists. contg. various ratios of 4% solns. of aces. and  
of sugars or of acids, and the initial osmotic rate (ml./  
min.) was plotted against the compn., the min. in the  
curve being  $x$ . The initial osmotic rate was detd. by ex-  
trapolating to zero time the osmotic rate vs. time curve.  
 $x$  was detd. for glucose, maltose, fructose, formic, acetic,  
propionic, oxalic, maleic, succinic, lactic, tartaric, citric,  
boric, and phosphoric acids with MeOH, EtOH, PrOH,  
sec-PrOH, and  $n$ -,  $sec$ -, and  $tert$ -BuOH. H. S.

G



CH

The structure of orthotelluric, phosphorous, and arsenious acids. B. Stehlik. *Collection Czechoslov. Chem. Commun.* 16, 241-7 (1949) (in English).—See C.A. 43, 8834gi.  
E. J. C.



Crystal structure of tellurium dioxide. B. Stehlik and  
J. Balak (Masaryk Univ., Brno). *Collection Czechoslov.  
Chem. Commun.* 14, 505-607 (1949) (in English). See -  
C.A. 43, 8781j. P. E. Brand

CA

Osmometric study of ammonium chloride. H. Stehlik  
(Slovak Tech. Univ., Bratislava). *Collection Cihlarova*.  
*Chem. Commun.* 10, 998-10 (1949) (in English). -- See  
C.A. 43, 8817g. ... N. M. S.

CA

2

Thermodynamic calculation of the temperature in the  
normal oxygen-acetylene flame. Blahoslav Stehlik  
(Slovak Tech. Univ., Bratislava, Czech.). *Chem.  
Zvesti* 6, 1-8(1960). See also 5410. *Method (Einde des  
flammas de soudure (C.A., 20, 2541'))* of calcn. is utilized.  
From the new data the calcn. for the temp. of the flame  
is:  $T = (3307 + 0.70t)^\circ\text{K}$ , where  $t$  is the initial temp.  
Jan Miska

CA 118

The molecular structure of glycine and  $\alpha$ -alanine.  
- Blahoslav Stehlik, Alexander Tkáč, and Nadě Lísková  
(Slovak Tech. Univ., Bratislava, Czech.). *Chem.  
Zvesti* 4, 53-9(1950).—The osmometric measurements,  
especially the indication of an anomalous hydrochloride,  
show that both glycine and  $\alpha$ -alanine do not have the  
form of the dipolar ion but the form of strong polar hydro-  
gens of the amino group. These measurements agree with  
spectroscopic ones but differ in their explanations. The  
crystal structure is discussed.  
Jan Mícha